

Standard Reduction Potentials of Some Biochemically Important Half-Reactions

Half-Reaction	E° (V)
$\frac{1}{2}\text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{H}_2\text{O}$	0.815
$\text{SO}_4^{2-} + 2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{SO}_3^{2-} + \text{H}_2\text{O}$	0.48
$\text{NO}_3^- + 2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{NO}_2^- + \text{H}_2\text{O}$	0.42
Cytochrome a_3 (Fe^{3+}) + $\text{e}^- \rightleftharpoons$ cytochrome a_3 (Fe^{2+})	0.385
$\text{O}_2(\text{g}) + 2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{H}_2\text{O}_2$	0.295
Cytochrome a (Fe^{3+}) + $\text{e}^- \rightleftharpoons$ cytochrome a (Fe^{2+})	0.29
Cytochrome c (Fe^{3+}) + $\text{e}^- \rightleftharpoons$ cytochrome c (Fe^{2+})	0.235
Cytochrome c_1 (Fe^{3+}) + $\text{e}^- \rightleftharpoons$ cytochrome c_1 (Fe^{2+})	0.22
Cytochrome b (Fe^{3+}) + $\text{e}^- \rightleftharpoons$ cytochrome b (Fe^{2+}) (mitochondrial)	0.077
Ubiquinone + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ ubiquinol	0.045
Fumarate $^-$ + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ succinate $^-$	0.031
FAD + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ FADH $_2$ (in flavoproteins)	~0
Oxaloacetate $^-$ + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ malate $^-$	-0.166
Peruvate $^-$ + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ lactate $^-$	-0.185
Acetaldehyde + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ ethanol	-0.197
FAD + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ FADH $_2$ (free coenzyme)	-0.219
$\text{S} + 2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{H}_2\text{S}$	-0.23
Lipoic acid + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ dihydrolipoic acid	-0.29
$\text{NAD}^+ + \text{H}^+ + 2\text{e}^- \rightleftharpoons \text{NADH}$	-0.315
$\text{NADP}^+ + \text{H}^+ + 2\text{e}^- \rightleftharpoons \text{NADPH}$	-0.320
Cystine + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ 2 cysteine	-0.340
Acetoacetate $^-$ + $2\text{H}^+ + 2\text{e}^- \rightleftharpoons$ β -hydroxybutyrate $^-$	-0.346
$\text{H}^+ + \text{e}^- \rightleftharpoons \frac{1}{2}\text{H}_2$	-0.421
Acetate $^-$ + $3\text{H}^+ + 2\text{e}^- \rightleftharpoons$ acetaldehyde + H_2O	-0.581

Source: Mostly from Loach, P.A., In Fasman, G.D. (Ed.), *Handbook of Biochemistry and Molecular Biology* (3rd ed.), Physical and Chemical Data, Vol. 1, pp. 123-130, CRC Press (1976).

This is Table 13-3 from Voet, Voet and Pratt *Fundamentals of Biochemistry* 2002, or 13-7 in Nelson and Cox, *Principles of Biochemistry* 2008